

Fractional Abundances

Find the weighted average atomic mass given the following fractional abundances:

1)

<u>Isotope</u>	<u>Mass (amu)</u>	<u>FA</u>
B-10	10.013	.1978
B-11	11.009	.8022

2)

<u>Isotope</u>	<u>Mass (amu)</u>	<u>FA</u>
O-16	15.995	.99761
O-17	16.999	.00039
O-18	17.999	.0020

3)

<u>Mass (amu)</u>	<u>FA</u>
38.964	.9326
39.964	.0001
40.962	.0673

What is the identity of this element?

4)

<u>Mass (amu)</u>	<u>FA</u>
69.924	.2123
71.922	.2766
72.923	.0773
73.921	.3594
75.921	.0744

What is the identity of this element?

Nuclear Properties

Using the periodic table, complete the following table.

Isotope Symbol	Element Name	Atomic Number	Mass Number	# of Protons	# of Electrons	# of Neutrons
${}^7\text{Li}$						
		7	15			
				19		21
					34	40
		44				60
${}^{119}\text{Sn}$						
		60	144			
			144	62		
			186		74	
		94				148
${}^{250}\text{Cf}$						